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The Mole and Volume Directions: For gases at STP (273K and 1.00 atm pressure), one mole occupies a volume of 22.4L. What volume will the following quantities of gasses occupy at STP? 1) 5.00 mole of H<sub>2</sub> ANSWER: 112. L 2) 2.50 moles of O<sub>2</sub> ANSWER: 56.0 L 3) 0.850 moles of N<sub>2</sub> ANSWER: 19.0 L 4) 3.75 moles of CO<sub>2</sub> ANSWER: 84.0 L 5) 0.70 mole of NH<sub>3</sub>

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MOLE-MOLE PROBLEMS N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub> How many moles of hydrogen are needed to completely react with two moles of nitrogen? 2. 2KClO<sub>3</sub> → 2KCl + 3O<sub>2</sub> How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate? 3. Zn + 2HCl → ZnCl<sub>2</sub> + H<sub>2</sub> How many moles of hydrogen are produced from the reaction of three moles

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